

$\text{AgF}(s) \longrightarrow \text{Ag}^+(aq) + \text{F}^-(aq)$	-15	to be calculated
$\text{Ag}^+(g) \longrightarrow \text{Ag}^+(aq)$	-464	

2 (a) Define the term enthalpy of hydration of an ion.

[2 marks]

Enthalpy change when 1 mole
of gaseous ions form aqueous
ions.

2 (b) Use data from Table 2 to calculate a value for the enthalpy of hydration of the chloride ion.

[2 marks]

$\Delta H_{\text{solution}}$

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