

Question 1

A 5.00 L sample of $\text{N}_2(\text{g})$ is held at a pressure of 2.00 atm and a temperature of 315 K. Assuming ideal-gas behavior, how many moles of $\text{N}_2(\text{g})$ are present in this sample?

- 2.69 mol
- 0.208 mol
- 0.387 mol
- 0.004 mol

Question 2

At a temperature of 500°C and a pressure of 699 Torr, sulfur vapor has a density of 3.71 g/dm^3 . What molecular formula for sulfur is compatible with this set of conditions?

- S_2
- S_8
- S_6
- $(\text{S}_{25})_2$

Question 3

Consider a gas for which the equation of state $P(V_m - b) = RT$ (with $b = 0.04 \text{ dm}^3/\text{mol}$) can be used to describe its P-V-T behavior. Suppose a 2-mol sample of this gas is compressed isothermally and reversibly from an initial volume of 10 dm^3 to a final volume of 1 dm^3 . How much work (w) is done on the gas sample in this compression process, and how much is the internal energy (U) of the gas changed?

- $w = 11.7 \text{ kJ}$ and $\Delta U = 11.7 \text{ kJ}$
- $w = 11.9 \text{ kJ}$ and $\Delta U = 0$
- $w = 5.85 \text{ kJ}$ and $\Delta U = -5.85 \text{ kJ}$
- $w = 5.85 \text{ kJ}$ and $\Delta U = 0$

Question 4

Consider a process in which a sample of water vapor expands reversibly and adiabatically from a state in which its pressure is 60 Torr and its volume is 0.5 dm^3 to a state in which its volume is 3 dm^3 . If it is assumed that the water vapor behaves as an ideal gas (with respect to its P-V-T properties) and that the ratio of its $C_{m,m}$ and $C_{m,v}$ heat-capacity quantities is given by $\gamma = C_{m,m}/C_{m,v} = 1.3$, what will be the final pressure of the vapor sample?

- 5.83 bar
- 230 Torr
- 5.84 Torr
- There is not enough information provided to answer this question.

Question 5

Consider a process in which exactly 107 J of heat is added to 1 mol of $\text{CH}_3\text{OH}(\text{g})$ under constant-volume conditions. If the initial temperature of the methanol vapor sample is 298 K, what will be the temperature of the vapor after the 107 J of heat is added? You may assume ideal-gas behavior and use a value of $43.89 \text{ J/K}\cdot\text{mol}$ for $C_{m,m}$ (the molar heat capacity of methanol vapor under constant-pressure conditions).

- 301.0 K
- 300.4 K
- 295.0 K
- None of the above answers is correct.

Question 6

If a person breathes at a rate of 14 respirations per minute, and each respiration involves 0.5 dm^3 of air, what will be the total mass of air breathed in a day? (Take $T = 300 \text{ K}$ and $P = 101.3 \text{ kPa}$, and assume that the air behaves as an ideal gas with a molar mass of 0.029 kg/mol .)

- 11.9 kg
- 119 g
- 238 g
- 14.6 kg

Question 7

If you have 1.55 mol of $\text{F}_2(\text{g})$ at 1.11 atm pressure and a temperature of 0°C , what is the volume of this gas sample, assuming ideal-gas behavior?

- 31.3 L
- 22.4 L
- 3170 L
- 0.0821 L

Question 8

The density of ice, $\text{H}_2\text{O}(\text{s})$, at 273 K, is 0.915 kg/dm^3 , and that of liquid water is 0.99987 kg/dm^3 . How much thermodynamic work must be done to melt 1 mol of ice at 1 bar pressure?

- 0.167 J
- 0.167 kJ
- 1.67 J
- 39.9 cal

Question 9

The heat capacity of a substance can be measured under either constant-volume or constant-pressure conditions. Which of the following sets of relationships between C_v (heat capacity under constant-volume conditions) and C_p (heat capacity under constant-pressure conditions) are always true?

- $C_p(\text{gas}) > C_v(\text{gas})$; $C_p(\text{solid}) = C_v(\text{solid})$; and $C_p(\text{gas}) >> C_p(\text{solid})$
- $C_v(\text{gas}) = C_v(\text{solid})$ and $C_p(\text{liquid}) = C_v(\text{liquid})$
- $C_v(\text{gas}) > C_v(\text{liquid}) > C_v(\text{solid})$ and $C_p > C_v$ in all phases
- C_p approaches a value of 0 as the sample temperature approaches 0 K but increases without bound as the sample temperature is increased.

Multiple Choice Physical Chemistry Exam Questions

S Marginson



Multiple Choice Physical Chemistry Exam Questions:

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Table of Contents Multiple Choice Physical Chemistry Exam Questions

1. Understanding the eBook Multiple Choice Physical Chemistry Exam Questions
 - The Rise of Digital Reading Multiple Choice Physical Chemistry Exam Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Physical Chemistry Exam Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Physical Chemistry Exam Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Physical Chemistry Exam Questions
 - Personalized Recommendations
 - Multiple Choice Physical Chemistry Exam Questions User Reviews and Ratings
 - Multiple Choice Physical Chemistry Exam Questions and Bestseller Lists
5. Accessing Multiple Choice Physical Chemistry Exam Questions Free and Paid eBooks
 - Multiple Choice Physical Chemistry Exam Questions Public Domain eBooks
 - Multiple Choice Physical Chemistry Exam Questions eBook Subscription Services
 - Multiple Choice Physical Chemistry Exam Questions Budget-Friendly Options
6. Navigating Multiple Choice Physical Chemistry Exam Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Physical Chemistry Exam Questions Compatibility with Devices
 - Multiple Choice Physical Chemistry Exam Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Physical Chemistry Exam Questions
 - Highlighting and Note-Taking Multiple Choice Physical Chemistry Exam Questions
 - Interactive Elements Multiple Choice Physical Chemistry Exam Questions

8. Staying Engaged with Multiple Choice Physical Chemistry Exam Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Physical Chemistry Exam Questions
9. Balancing eBooks and Physical Books Multiple Choice Physical Chemistry Exam Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Physical Chemistry Exam Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Physical Chemistry Exam Questions
 - Setting Reading Goals Multiple Choice Physical Chemistry Exam Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Physical Chemistry Exam Questions
 - Fact-Checking eBook Content of Multiple Choice Physical Chemistry Exam Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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