

Gas Stoichiometry: Multiple Choice

1. What volume of pure oxygen gas, O_2 , measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0.200 mol Ag_2O according to the following equation:



- A) 1.12 L B) 4.48 L C) 2.24 L D) 8.96 L E) 17.82 L

ANS: B

1. How many moles of $KClO_3$ are needed to form 2.8 L of O_2 , measured at STP, according to the following reaction: $2KClO_3 \rightarrow 2KCl + 3O_2$

- A) 0.083 mol B) 0.25 mol C) 0.17 mol D) 0.50 mol E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101.325 kPa, is required to react completely with 5.61 g of KOH in the following reaction: $Cl_{2(g)} + 2KOH \rightarrow KCl + KClO + H_{2O}$

- A) 0.56 L B) 4.48 L C) 1.12 L D) 22.4 L E) 2.24 L

ANS: C

1. Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas: $NaH(s) + H_2O(l) \rightarrow NaOH(aq) + H_2(g)$

A sample of NaH weighing _____ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

- A) 2.93 B) 0.960 C) 0.925 D) 0.0388 E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to form water? $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$

- A) 1.8 L B) 3.6 L C) 2.0 L D) 2.4 L

ANS: A

1. The molar volume of a gas at STP is _____ L.

- A) 0.08206 B) 62.36 C) 1.00 D) 22.4 E) 14.7

ANS: D

1. If 0.5 L of $O_2(g)$ reacts with H_2 to produce 1 L of $H_2O(g)$, what is the volume of $H_2O(g)$ obtained from 1 L of $O_{2(g)}$?

- A) 0.5 L B) 1.5 L C) 2 L D) 2.5 L

ANS: C

Multiple Choice Questions On Chemical Stoichiometry

Michael Brown



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