

## Gas Stoichiometry: Multiple Choice

1. What volume of pure oxygen gas,  $O_2$ , measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0.200 mol  $Ag_2O$  according to the following equation:



- A) 1.12 L      B) 4.48 L      C) 2.24 L      D) 8.96 L      E) 17.82 L

ANS: B

1. How many moles of  $KClO_3$  are needed to form 2.8 L of  $O_2$ , measured at STP, according to the following reaction:  $2KClO_3 \rightarrow 2KCl + 3O_2$

- A) 0.083 mol    B) 0.25 mol    C) 0.17 mol    D) 0.50 mol    E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101.325 kPa, is required to react completely with 5.61 g of KOH in the following reaction:  $Cl_{2(g)} + 2KOH \rightarrow KCl + KClO + H_{2O}$

- A) 0.56 L      B) 4.48 L      C) 1.12 L      D) 22.4 L      E) 2.24 L

ANS: C

1. Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas:  $NaH(s) + H_2O(l) \rightarrow NaOH(aq) + H_2(g)$

A sample of  $NaH$  weighing \_\_\_\_\_ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

- A) 2.93      B) 0.960      C) 0.925      D) 0.0388      E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to form water?  $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$

- A) 1.8 L      B) 3.6 L      C) 2.0 L      D) 2.4 L

ANS: A

1. The molar volume of a gas at STP is \_\_\_\_\_ L.

- A) 0.08206    B) 62.36    C) 1.00    D) 22.4    E) 14.7

ANS: D

1. If 0.5 L of  $O_2(g)$  reacts with  $H_2$  to produce 1 L of  $H_2O(g)$ , what is the volume of  $H_2O(g)$  obtained from 1 L of  $O_{2(g)}$ ?

- A) 0.5 L      B) 1.5 L      C) 2 L      D) 2.5 L

ANS: C

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