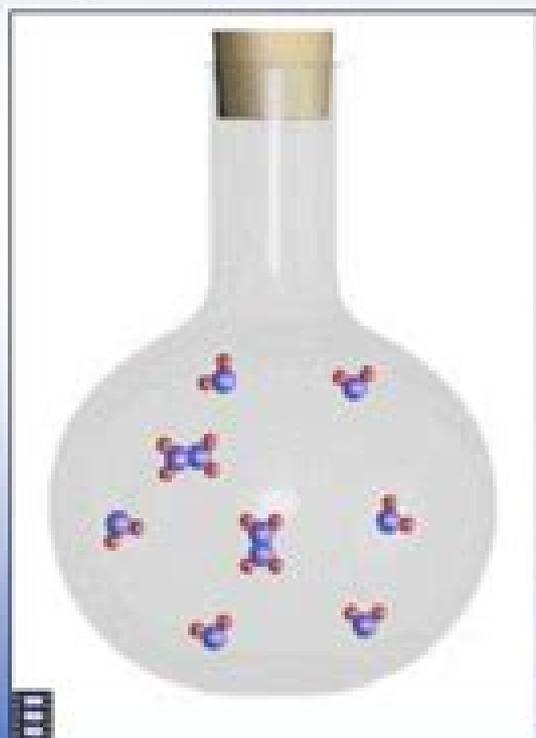


Nitrogen Dioxide Equilibrium

Increase P in the system by reducing the volume (at constant T).



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]} = 0.0058 \text{ at } 298 \text{ K}$$

In gaseous system the equilibrium will shift to the side with **fewer molecules** (in order to reduce the P).

Therefore, reaction shifts _____
and P of NO_2 _____ and P of
 N_2O_4 _____.

Equilibrium Nitrogen Dioxide Lab

AW Chickering



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