

16 • Chemical Equilibrium**STATION 1 • MASS ACTION EXPRESSIONS**

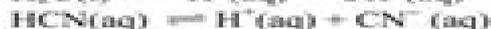
Write the mass action expression for the equilibrium: $\text{Fe}_3\text{O}_4(\text{s}) + 4 \text{H}_2(\text{g}) \rightleftharpoons 3 \text{Fe}(\text{s}) + 4 \text{H}_2\text{O}(\text{g})$

Write the mass action expression for the equilibrium: $\text{Ca}(\text{OH})_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2 \text{OH}^-(\text{aq})$

The equilibrium constant, $K_c = 7.9 \times 10^{-6}$, is this equilibrium system reactant or product-favored. _____

16 • Chemical Equilibrium**STATION 2 • MANIPULATING K**

Given:



$$K_c = 1 \times 10^{-14}$$

$$K_c = 4.0 \times 10^{-10}$$

Calculate K_c for this reaction:



$$K_c = ???$$

This reaction is _____-favored.

Equilibrium Worksheet Pasadena

William Guenther



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Scientific and Technical Aerospace Reports ,1980 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database **The New Adapted Physical Education** Janet A. Seaman, Karen P. DePauw, 1989

Government Reports Announcements & Index ,1988-12 INIS Atomindex ,1980 **Chemical Equilibrium** William Guenther, 1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

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