

CHAPTER 7

STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) _____.

Such an attachment may form by the attraction of the positively charged (2) _____ of one atom for the negatively charged (3) _____ of another atom, or by the attraction of charged atoms, which are called (4) _____. The attractions may also involve (5) _____ electrons, which are the electrons in the outermost (6) _____. The (7) _____ are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) _____. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

- _____ 10. A positively charged ion is called an anion.
- _____ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- _____ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- _____ 13. A sodium atom tends to lose one electron when it reacts.
- _____ 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- _____ 15. A Cl^- ion is an example of a cation.
- _____ 16. The ending *-ide* is used to designate an anion.
- _____ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Ionic Compounds Study Guide

Ying-Ying Zheng



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Table of Contents Ionic Compounds Study Guide

1. Understanding the eBook Ionic Compounds Study Guide
 - The Rise of Digital Reading Ionic Compounds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compounds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Compounds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compounds Study Guide
 - Personalized Recommendations
 - Ionic Compounds Study Guide User Reviews and Ratings
 - Ionic Compounds Study Guide and Bestseller Lists
5. Accessing Ionic Compounds Study Guide Free and Paid eBooks
 - Ionic Compounds Study Guide Public Domain eBooks
 - Ionic Compounds Study Guide eBook Subscription Services
 - Ionic Compounds Study Guide Budget-Friendly Options

6. Navigating Ionic Compounds Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Compounds Study Guide Compatibility with Devices
 - Ionic Compounds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Compounds Study Guide
 - Highlighting and Note-Taking Ionic Compounds Study Guide
 - Interactive Elements Ionic Compounds Study Guide
8. Staying Engaged with Ionic Compounds Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Compounds Study Guide
9. Balancing eBooks and Physical Books Ionic Compounds Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Compounds Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Compounds Study Guide
 - Setting Reading Goals Ionic Compounds Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Compounds Study Guide
 - Fact-Checking eBook Content of Ionic Compounds Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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