

# CHEMICAL BONDS

## Ionic Bonding

Metals and non-metals react together to form \_\_\_\_\_ compounds. They are held together by ionic bonds. Ionic bonding is a force of \_\_\_\_\_ attraction between two \_\_\_\_\_ atoms (called \_\_\_\_\_). An atom becomes charged by either donating or accepting \_\_\_\_\_ electrons from another atom. A valence electron is one from the \_\_\_\_\_ shell of the atom.

Since opposites attract, there is an attraction between the positively charged and negatively charged ions that holds them together.



If an atom \_\_\_\_\_ an electron, it becomes positively charged. This is then called a \_\_\_\_\_. \_\_\_\_\_ always donate electrons. If an atom accepts an electron, it becomes \_\_\_\_\_. This is called an \_\_\_\_\_. Non-metals always accept electrons.

Atoms do this in order to become '\_\_\_\_\_'. A stable atom is one which has obtained a \_\_\_\_\_ of electrons.

*How do we know which atoms donate and which accept electrons?*

Sodium has \_\_\_\_\_ lone electron in the outer shell. It is easiest for this atom to \_\_\_\_\_ one electron than to try and find seven more to fill the outer shell.



All \_\_\_\_\_ donate electrons and therefore become cations.

All \_\_\_\_\_ accept electrons and therefore become anions.

Chlorine has \_\_\_\_\_ electrons in its outer shell. Therefore, it is easiest for Chlorine to \_\_\_\_\_ one electron than to \_\_\_\_\_ to try and donate seven.



# Ionic And Covalent Bonding Study Guide

**Kenneth J. Hughes, Paul B. Kelter**



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