

CHAPTER 7

STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

| | | | | |
|---------------|-----------|-----------------------------|---------|-------------|
| chemical bond | electrons | energy level | ions | noble gases |
| nucleus | octet | pseudo-noble gas formations | valence | |

The force that holds two atoms together is called a(n) (1) _____.
 Such an attachment may form by the attraction of the positively charged
 (2) _____ of one atom for the negatively charged
 (3) _____ of another atom, or by the attraction of charged atoms, which are called
 (4) _____. The attractions may also involve
 (5) _____ electrons, which are the electrons in the outermost
 (6) _____. The (7) _____ are a family of elements that have very
 little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a
 stable (8) _____. The relatively stable electron structures developed by loss of
 electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

- _____ 10. A positively charged ion is called an anion.
- _____ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- _____ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- _____ 13. A sodium atom tends to lose one electron when it reacts.
- _____ 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- _____ 15. A Cl^- ion is an example of a cation.
- _____ 16. The ending *-ide* is used to designate an anion.
- _____ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Ions And Ionic Compounds Study Guide

Michael Seilmaier



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