

Section 7 - Radioactivity
 BY HB SCIENCE RESOURCES

Introduction
 Radioactivity is the spontaneous emission of particles or electromagnetic radiation from an unstable nucleus.

Types of Radioactivity

- Alpha (α)**: Consists of two protons and two neutrons. It is the least penetrating and is stopped by a sheet of paper.
- Beta (β)**: Consists of a high-speed electron. It is more penetrating than alpha and is stopped by a thin metal plate.
- Gamma (γ)**: Consists of high-energy electromagnetic waves. It is the most penetrating and is stopped by a thick lead or concrete wall.

Half-life
 The half-life of a radioactive substance is the time taken for half of the original substance to decay.

Radioactive Decay
 The number of undecayed nuclei in a sample of a radioactive substance decreases exponentially over time.

Radioactive Dating
 The age of a sample can be determined by measuring the ratio of the parent isotope to the daughter product.

Background Radiation
 There is always some background radiation present in the environment.

Uses of Radioactivity
 Radioactive isotopes are used in medicine, industry, and agriculture.

Diagram 1: Alpha Decay
 ${}^A_ZX \rightarrow {}^{A-4}_{Z-2}Y + {}^4_2\text{He}$

Diagram 2: Beta Decay
 ${}^A_ZX \rightarrow {}^A_{Z+1}Y + e^- + \bar{\nu}_e$

Diagram 3: Gamma Decay
 ${}^A_ZX^* \rightarrow {}^A_ZX + \gamma$

Table 1: Properties of Radioactive Particles

Particle	Symbol	Charge	Penetrating Power
Alpha	${}^4_2\text{He}$	+2	Low
Beta	e^-	-1	Medium
Gamma	γ	0	High

Table 2: Radioactive Isotopes

Isotope	Half-life	Use
${}^{14}_6\text{C}$	5730 years	Dating of organic remains
${}^{238}_{92}\text{U}$	4.5 billion years	Energy production
${}^{60}_{27}\text{Co}$	5.27 years	Medical treatment

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Karin Nielsen-Saines



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