balance nuclear reaction equations
aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

$$^{26}_{13}Al \rightarrow ^{0}_{1}e + ^{26}_{12}Mg$$
$$^{40}_{19}K \rightarrow ^{0}_{-1}e + ^{40}_{20}Ca$$
$$^{238}_{92}U \rightarrow ^{4}_{2}He + ^{234}_{90}Th$$

electron configurations
Write the electron configuration for C.
Write the electron configuration for Cr.
Write the shorthand notation for Bi.

$1s^2 2s^2 2p^2$
$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^5$
$[Xe] 6s^2 4f^{14} 5d^{10} 6p^3$

NOMENCLATURE [naming compounds, writing formulas]
Write the formula for each of the following compounds:
a. Sulfur hexafluoride — SF_6
b. Lithium nitride — Li_3N
c. Chromium (III) carbonate — $Cr_2(CO_3)_3$
d. Tin (II) chloride — $SnCl_2$
e. Ammonium acetate — $NH_4C_2H_3O_2$ (or CH_3COONH_4)
f. Mercury (I) chloride — Hg_2Cl_2
g. Potassium bromate — $KBrO_3$
h. Hydrosulfuric acid — $H_2S_{(aq)}$
i. Chloric acid — $HClO_3$
j. Sulfurous acid — H_2SO_3

Name each of the following compounds.
a. $CuSO_4$ — Copper (II) sulfate
b. AlF_3 — aluminum fluoride
c. HI — hydrogen iodide
d. NO — nitrogen monoxide
e. $H_2Se_{(aq)}$ — hydroselenic acid
f. HNO_3 — nitric acid
g. HNO_2 — nitrous acid
h. $NaHSO_4$ — Sodium hydrogen sulfate

Half-life

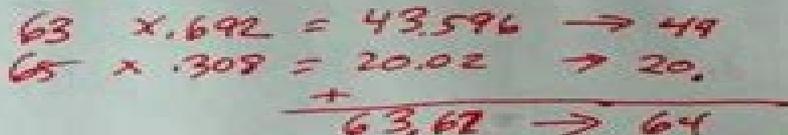Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g? $40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$ 1 2 3 4 5 → 61.5 yo

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?
3 half lives $100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5mg$ 1 2 3

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$63 \times .692 = 43.596 \rightarrow 44$
$65 \times .308 = 20.02 \rightarrow 20.$
$\underline{+}$
$63.62 \rightarrow 64$

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Ensheng Dong

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