

# ACS GENERAL CHEMISTRY FINAL STUDY GUIDE

Avagadro's Number -  $N_A = 6.023 \times 10^{23}$

Percent Yield -  $(\text{Actual Yield}/\text{Theoretical Yield}) \times 100$

Percent Composition (of an Element) -  $[(\text{moles of element} \times \text{molar mass of element})/(\text{molar mass of compound})] \times 100$

Dilution Equation -  $M_1V_1 = M_2V_2$   
Mass and Volume

Boyle's Law -  $P_1V_1 = P_2V_2$   
Pressure and Volume

Charles' Law -  $V_1/T_1 = V_2/T_2$   
Volume and Temperature

Ideal Gas Law -  $PV = nRT$

Volume of an Ideal Gas at STP - 22.4 L

Density (Ideal Gas) -  $(P \times \text{Molar Mass})/(RT)$

Definition of Entahipy -  $H = E + PV$

Definition of Heat Capacity -  $C = ms$   
 $s = \text{specific heat}$

ion - different number of electrons

isotopes - different number of neutrons

Periodic Table Trends - As you go down or left atomic radius increases  
As you go up or right ionization energy and electron affinity increases

Ionic Bonds - to charged pieces come together and then there are cations or anions

Covalent Bonds - are when two non metals share electrons

A Compound is probably soluble if it contains these anions: -  $\text{Cl}^-$   
 $\text{Br}^-$   
 $\text{I}^-$

# General Chemistry Final Study Guide

**LL Leslie**



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