

Genetics Study Guide

6 pages + 6-page Answer KEY

Biology

GENETICS STUDY GUIDE

1. The study of heredity is known as _____.
2. The study of heredity is known as _____.
3. What are the two main types of inheritance?
 - a. _____
 - b. _____
4. What are the two main types of inheritance?
 - a. _____
 - b. _____
5. What are the two main types of inheritance?
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60. What are the two main types of inheritance?
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ANSWER KEY INCLUDED

Genetics Study Guide Answer Key

F Rizvi



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Separate chapter on genetics focuses on the practical application to nursing care of patients Genetics in Clinical Practice
boxes address key topics such as genetic testing Alzheimer s disease sickle cell disease and genetics related ethical issues
Genetic Risk Alerts and Genetic Link headings highlight specific genetic issues related to body system assessments and
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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8 3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S}$

mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137}{149} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x $\frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 81.2 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{26.982 \text{ g Al}_2\text{O}_3}{81.2 \text{ g Al}_2\text{O}_3} \times 100 = 33.2\%$

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