

3. Attempt only 8 questions**8 X 2 = 16****A. What is activation of catalyst?****Activation of catalyst:**

A substance which promotes the activity of a catalyst is called a promoter or activator or activation of catalyst.

Example:

Hydrogenation of vegetable oil is carried out by Ni. The catalytic activity of Ni can be increased by using Cu and tellurium.

B. Differentiate between negative catalyst and auto-catalyst?

Sl. #	Negative catalyst	Auto catalyst
1	A substance which retards the rate of a reaction is called a negative catalyst. And this phenomenon is called negative catalysis.	In some reactions, a product formed acts as catalyst. And this phenomenon is called auto catalysis.

iii. Write four characteristic of enzyme catalysis?**Characteristic of enzyme catalysis:**

- 1) Enzymes are the most efficient catalysts known and they lower the energy of activation of a reaction.
- 2) Enzymes catalysis are highly specific. E.g., Urease catalyses the hydrolysis of urea only and it cannot hydrolyse any other amide even methyl urea.
- 3) Enzyme catalytic reactions have maximum rates at an optimum temperature.
- 4) The pH of the medium also controls the rates of enzyme catalytic reactions. Each enzyme has an optimum pH.

iv. Write four characteristics of a catalyst?**Characteristics of a catalyst:**

- 1) A catalyst cannot affect the equilibrium constant of a reaction. It only decreases the time to reach equilibrium.
- 2) A catalyst is specific in its reaction. A particular catalyst works for one reaction. It may not work for any other reaction.
- 3) The mechanism of a catalysed reaction is different from that of an uncatalysed reaction.
- 4) A catalyst cannot start a reaction, which is not thermodynamically feasible.

v. What is energy of activation?**Activation energy:**

The minimum amount of energy in addition to average K.E which the particles must have for the effective collisions is called activation energy.

$$E_a = -2.303 R \times \text{slope}$$

Activation energy provides information about the path a reaction occurs.

vi. Draw Arrhenius plot to calculate the energy of activation.

See the topic of Arrhenius equation.

vii. Name the factors which affect the rate of chemical reaction.

Following factors affect the rate of reaction.

- Nature of a reactant.
- Concentration of reactants.
- Surface area.
- Light.
- Temperature.
- Catalyst.

viii. What is the effect of concentration of reactants on the rate of reaction?

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WJ Hussar



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