

## Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
  - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
  - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
  - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
  - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
  - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
  - The type of sublevel being filled in successive elements of that block.
  - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
  - The Group 1 elements, except Hydrogen.
  - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
  - The Group 2 elements.
  - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
  - Transition Metals (d-block)
    - Group 3:  $ns^2 (n-1)d^1$
    - Group 4:  $ns^2 (n-1)d^2$
    - Group 5:  $ns^2 (n-1)d^3$
    - Group 6:  $ns^2 (n-1)d^4$  {actually is  $ns^1 (n-1)d^5$  for Cr & Mo}
    - Group 7:  $ns^2 (n-1)d^5$
    - Group 8:  $ns^2 (n-1)d^6$
    - Group 9:  $ns^2 (n-1)d^7$
    - Group 10:  $ns^2 (n-1)d^8$
    - Group 11:  $ns^2 (n-1)d^9$  {actually is  $ns^1 (n-1)d^{10}$  for all group 11 elements}
    - Group 12:  $ns^2 (n-1)d^{10}$
  - The group number is the sum of the s & d electrons.
- The transition elements
- p-block
  - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
  - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
  - The Group 17 elements.

# Modern Chemistry Chapter 5 Review Answer Key

**DJ Losen**



**Modern Chemistry Chapter 5 Review Answer Key:**

## Whispering the Techniques of Language: An Psychological Journey through **Modern Chemistry Chapter 5 Review Answer Key**

In a digitally-driven earth where displays reign great and immediate conversation drowns out the subtleties of language, the profound secrets and emotional nuances concealed within phrases frequently move unheard. However, set within the pages of **Modern Chemistry Chapter 5 Review Answer Key** a fascinating fictional prize blinking with natural emotions, lies an exceptional quest waiting to be undertaken. Penned by an experienced wordsmith, that charming opus invites viewers on an introspective trip, lightly unraveling the veiled truths and profound affect resonating within ab muscles fabric of every word. Within the psychological depths of this touching evaluation, we shall embark upon a genuine exploration of the book is core themes, dissect its interesting writing fashion, and fail to the strong resonance it evokes strong within the recesses of readers hearts.

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