

# Answer Key – Class 10 Science: Chemical Reactions and Equations

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## Section A: Multiple Choice Questions

1. c)  $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
2. b) Copper is deposited on iron
3. a) Hydrogen
4. d) All of the above
5. d)  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$

## Section B: Very Short Answer Questions

6.  $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2 \uparrow$
7. A precipitation reaction is a reaction where an insoluble solid (precipitate) is formed.  
Example:  
 $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} \downarrow + \text{NaNO}_3$
8. Colour changes from blue to light green. Equation:  $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
9. Double displacement reaction.  
Balanced:  $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 \downarrow + 2\text{NaCl}$
10. (i) Change in colour  
(ii) Evolution of gas / Formation of precipitate

## Section C: Short Answer Questions

11. A redox reaction involves both oxidation and reduction.  
Example:  $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$   
Zn is oxidised (reducing agent),  $\text{Cu}^{2+}$  is reduced (oxidising agent).
12. On heating, lead nitrate decomposes to give lead oxide, nitrogen dioxide and oxygen.  
 $2\text{Pb(NO}_3)_2 \rightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$   
Type: Thermal decomposition reaction.

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