

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter in class.

Please use this material, but make sure to review other notes, activities, handouts, worksheets, etc.

If you already take advanced-level AP/IB class, your school is listed for state or PA-IB. Use that.

**** Use items with the words, examples, and/or application of this material. ****

COURSES:

- College-level (with general or interdisciplinary) courses: GC 42
- Community-college-level or transfer courses: GC 43
- 5.1: Describe the organization of the modern Periodic Table; distinguish between periods and groups. IT 41 42
- 5.2: Use the Periodic Table to identify elements about their properties (chemical and physical), reactivity, energy, and atomic size, of that elements. IT 41 42
- 5.3: Identify common compounds, molecules, and ionic gases; identify the representative groups and/or main groups (groups 1-2, 13, 14, 15, 16, 17, 18) IT 41 42
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity. IT 41 42

5.1: FACTS OF THE PERIODIC TABLE

• Introductory

- periodic law - the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements.
- periodic table - an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- horizontalities - a number of the same number of elements, known as the number, range from 2 (hydrogen) to 18 (barium).
- verticality - set of the same of heavy metals that elements that decrease from hydrogen (atomic number 1) through barium (atomic number 138) in the periodic table.

• Chapter Highlights

- The periodic law states that the physical and chemical properties of the elements and periodic functions of their atomic numbers.
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- shell model - one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- shellness model - one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- representative elements - one of the metals that can use the inner shell electrons using the outer shell to bond.
- transition elements - an element in the d-block or f-block of the periodic table.
- halogens - one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts.

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