

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- |                             |   |
|-----------------------------|---|
| a. Sulfur hexafluoride      | $\text{SF}_6$   |
| b. Lithium nitride          | $\text{Li}_3\text{N}$   |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$  |
| d. Tin (II) chloride        | $\text{SnCl}_2$   |
| e. Ammonium acetate         | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$ ) |
| f. Mercury (I) chloride     | $\text{Hg}_2\text{Cl}_2$  |
| g. Potassium bromate        | $\text{KBrO}_3$   |
| h. Hydrosulfuric acid       | $\text{H}_2\text{S(aq)}$  |
| i. Chloric acid             | $\text{HClO}_3$   |
| j. Sulfurous acid           | $\text{H}_2\text{SO}_3$   |

Name each of the following compounds:

- |                              |                         |
|------------------------------|-------------------------|
| a. $\text{CuSO}_4$           | Copper (II) sulfate     |
| b. $\text{AlF}_3$            | aluminum fluoride       |
| c. HI                        | hydrogen iodide         |
| d. NO                        | nitrogen monoxide       |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid       |
| f. $\text{HNO}_3$            | nitric acid             |
| g. $\text{HNO}_2$            | nitrous acid            |
| h. $\text{NaHSO}_4$          | sodium hydrogen sulfate |

### Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1            2            3            4            5

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \text{ mg}$$

1            2            3

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 44 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

# Modern Chemistry Homework Section 7 Section Answers

**DP Hallahan**



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## **Table of Contents Modern Chemistry Homework Section 7 Section Answers**

1. Understanding the eBook Modern Chemistry Homework Section 7 Section Answers
  - The Rise of Digital Reading Modern Chemistry Homework Section 7 Section Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Homework Section 7 Section Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Homework Section 7 Section Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Homework Section 7 Section Answers
  - Personalized Recommendations
  - Modern Chemistry Homework Section 7 Section Answers User Reviews and Ratings

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- Modern Chemistry Homework Section 7 Section Answers and Bestseller Lists
- 5. Accessing Modern Chemistry Homework Section 7 Section Answers Free and Paid eBooks
  - Modern Chemistry Homework Section 7 Section Answers Public Domain eBooks
  - Modern Chemistry Homework Section 7 Section Answers eBook Subscription Services
  - Modern Chemistry Homework Section 7 Section Answers Budget-Friendly Options
- 6. Navigating Modern Chemistry Homework Section 7 Section Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Homework Section 7 Section Answers Compatibility with Devices
  - Modern Chemistry Homework Section 7 Section Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Homework Section 7 Section Answers
  - Highlighting and Note-Taking Modern Chemistry Homework Section 7 Section Answers
  - Interactive Elements Modern Chemistry Homework Section 7 Section Answers
- 8. Staying Engaged with Modern Chemistry Homework Section 7 Section Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Homework Section 7 Section Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Homework Section 7 Section Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Homework Section 7 Section Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Homework Section 7 Section Answers
  - Setting Reading Goals Modern Chemistry Homework Section 7 Section Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Homework Section 7 Section Answers
  - Fact-Checking eBook Content of Modern Chemistry Homework Section 7 Section Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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