

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol, C₂H₅OH, are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL C}_2\text{H}_5\text{OH}}{1000 \text{ mL C}_2\text{H}_5\text{OH}} \times \frac{1 \text{ L C}_2\text{H}_5\text{OH}}{1 \text{ L C}_2\text{H}_5\text{OH}} \times \frac{1.5 \text{ mole C}_2\text{H}_5\text{OH}}{1 \text{ L C}_2\text{H}_5\text{OH}} = 0.098 \text{ moles C}_2\text{H}_5\text{OH}$$

2. How many liters of a 6.0 M solution of acetic acid CH₃COOH, contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles CH}_3\text{COOH}}{6.0 \text{ moles CH}_3\text{COOH}} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M H₂SO₄ = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M Ca(OH)₂ = 1 N

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