

# The Mole Study Guide

The Mole

**Molar Mass**

1. Calculate the molar mass of the following compounds:

- (a)  $\text{H}_2\text{O}$
- (b)  $\text{CO}_2$
- (c)  $\text{NaCl}$
- (d)  $\text{CH}_4$
- (e)  $\text{C}_6\text{H}_{12}\text{O}_6$

**Molar Volume**

2. Calculate the number of moles in the following volumes of gas at STP:

- (a) 22.4 L
- (b) 11.2 L
- (c) 5.6 L
- (d) 2.24 L

**Molar Concentration**

3. Calculate the molar concentration of the following solutions:

- (a) 0.5 mol of  $\text{NaCl}$  in 2.0 L of solution
- (b) 1.0 mol of  $\text{H}_2\text{SO}_4$  in 0.5 L of solution
- (c) 2.0 mol of  $\text{KNO}_3$  in 1.0 L of solution

4. Calculate the volume of gas at STP produced from the following reactions:

- (a) 2.0 mol of  $\text{Zn}$  reacts with  $\text{HCl}$
- (b) 1.0 mol of  $\text{CaCO}_3$  reacts with  $\text{HCl}$

5. Calculate the mass of the following substances:

- (a) 2.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

6. Calculate the number of particles in the following amounts:

- (a) 1.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

7. Calculate the number of moles in the following amounts:

- (a)  $6.02 \times 10^{23}$  particles of  $\text{H}_2\text{O}$
- (b)  $1.12 \times 10^{24}$  particles of  $\text{CO}_2$
- (c)  $3.01 \times 10^{23}$  particles of  $\text{NaCl}$

8. Calculate the volume of gas at STP produced from the following reactions:

- (a) 2.0 mol of  $\text{Zn}$  reacts with  $\text{HCl}$
- (b) 1.0 mol of  $\text{CaCO}_3$  reacts with  $\text{HCl}$

9. Calculate the mass of the following substances:

- (a) 2.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

10. Calculate the number of particles in the following amounts:

- (a) 1.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

11. Calculate the number of moles in the following amounts:

- (a)  $6.02 \times 10^{23}$  particles of  $\text{H}_2\text{O}$
- (b)  $1.12 \times 10^{24}$  particles of  $\text{CO}_2$
- (c)  $3.01 \times 10^{23}$  particles of  $\text{NaCl}$

12. Calculate the volume of gas at STP produced from the following reactions:

- (a) 2.0 mol of  $\text{Zn}$  reacts with  $\text{HCl}$
- (b) 1.0 mol of  $\text{CaCO}_3$  reacts with  $\text{HCl}$

13. Calculate the mass of the following substances:

- (a) 2.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

14. Calculate the number of particles in the following amounts:

- (a) 1.0 mol of  $\text{H}_2\text{O}$
- (b) 0.5 mol of  $\text{CO}_2$
- (c) 1.0 mol of  $\text{NaCl}$

15. Calculate the number of moles in the following amounts:

- (a)  $6.02 \times 10^{23}$  particles of  $\text{H}_2\text{O}$
- (b)  $1.12 \times 10^{24}$  particles of  $\text{CO}_2$
- (c)  $3.01 \times 10^{23}$  particles of  $\text{NaCl}$

# Mole Unit Study Guide Answers

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*Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry* Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface     *Student Study Guide with Programmed Problems to Accompany Fundamentals of Physics & Physics, Parts I & II* Stanley A. Williams, Kenneth Brownstein, Robert Gray, 1974

*Study Guide for E. Russell Hardwick's Introduction to Chemistry* Joan Bouillon, 1984     **Student Study Guide & Selected Solutions Manual** David D. Reid, 2007     **Chemistry, Study Guide**

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**Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 **Study Guide, Student Solutions Manual** Raymond A. Serway, John R. Gordon, 1998

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