1/5/16

Mole Ratio Worksheet

Key

1) Given this equation: $N_2 + 3\,H_2 \rightarrow 2\,NH_3$, write the following molar ratios:

a) N_2 / H_2 1:3

b) N_2 / NH_3 1:2

c) H_2 / NH_3 3:2

2) Given the following equation: $8\,H_2 + S_8 \rightarrow 8\,H_2S$, write the following molar ratios:

a) H_2 / H_2S 8:8 → 1:1 (reduced)

b) H_2 / S_8 8:1

c) H_2S / S_8 8:1

3) Answer the following questions for this equation: $2\,H_2 + O_2 \rightarrow 2\,H_2O$

a) What is the H_2 / H_2O molar ratio? 2:2 → 1:1 (reduced)

b) Suppose you had 20 moles of H_2 on hand and plenty of O_2, how many moles of H_2O could you make? $\frac{20\text{ mole } H_2 \mid 2\,H_2O}{2\,H_2} = 20\text{ mole } H_2O$

c) What is the O_2 / H_2O molar ratio?

1:2

d) Suppose you had 20 moles of O_2 and enough H_2, how many moles of H_2O could you make? $\frac{20\text{ mol } O_2 \mid 2\,H_2O}{1\,O_2} = 40\text{ mol } H_2O$

4) Use this equation: $N_2 + 3\,H_2 \rightarrow 2\,NH_3$, for the following problems

a) If you used 1 mole of N_2, how many moles of NH_3 could be produced? 2

b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required? $\frac{10\text{ mol } NH_3 \mid 1\,N_2}{2\,NH_3} = 5\text{ mol } N_2$

c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made? $\frac{3\text{ mol } H_2 \mid 2\,NH_3}{3\,H_2} = 2\text{ mol } NH_3$

d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required? $\frac{0.6\text{ mol } NH_3 \mid 3\,H_2}{2\,NH_3} = 0.9\text{ mol } H_2$

Mole Ratio Packet Answers

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