

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- | | | | |
|------------------------------|------------------------------------|----------------------------|------------------------------------|
| 1. copper (II) oxide | <u>CuO</u> | 11. manganese (III) oxide | <u>Mn₂O₃</u> |
| 2. mercury (I) oxide | <u>Hg₂O</u> | 12. vanadium (II) bromide | <u>VBr₂</u> |
| 3. gold (III) chloride | <u>AuCl₃</u> | 13. niobium(V) oxide | <u>Nb₂O₅</u> |
| 4. thallium (III) bromide | <u>TlBr₃</u> | 14. titanium (III) oxide | <u>Ti₂O₃</u> |
| 5. bismuth (V) oxide | <u>Bi₂O₅</u> | 15. titanium (III) nitride | <u>TiN</u> |
| 6. terbium (III) oxide | <u>Tb₂O₃</u> | 16. iron (II) oxide | <u>FeO</u> |
| 7. uranium (VI) oxide | <u>UO₃</u> | 17. cobalt (II) phosphide | <u>Co₃P₂</u> |
| 8. protactinium (IV) bromide | <u>PaBr₄</u> | 18. tin (II) oxide | <u>SnO</u> |
| 9. cerium (III) oxide | <u>Ce₂O₃</u> | 19. thulium (II) bromide | <u>TmBr₂</u> |
| 10. arsenic (V) sulphide | <u>As₂S₅</u> | 20. copper (II) bromide | <u>CuBr₂</u> |

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- | | | | |
|-----------------------------------|-----------------------------|------------------------------------|--------------------------------|
| 1. SnCl ₄ | <u>Tin (IV) Chloride</u> | 9. PdF ₄ | <u>Palladium (IV) Fluoride</u> |
| 2. BiBr ₅ | <u>Bismuth (V) Bromide</u> | 10. Os ₂ O ₃ | <u>Osmium (III) Oxide</u> |
| 3. PoO ₂ | <u>Polonium (IV) Oxide</u> | 11. MoBr ₂ | <u>Molybdenum (II) Bromide</u> |
| 4. PbI ₂ | <u>Lead (II) Iodide</u> | 12. VCl ₅ | <u>Vanadium (V) Chloride</u> |
| 5. HgO | <u>Mercury (II) Oxide</u> | 13. Mn ₂ O ₃ | <u>Manganese (III) Oxide</u> |
| 6. HgCl | <u>Mercury (I) Chloride</u> | 14. CoO | <u>Cobalt (II) Oxide</u> |
| 7. Au ₂ O ₃ | <u>Gold (III) Oxide</u> | 15. Np ₂ O ₃ | <u>Neptunium (III) Oxide</u> |
| 8. FeCl ₂ | <u>Iron (II) Chloride</u> | 16. V ₂ O ₅ | <u>Vanadium (V) Oxide</u> |

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Katrin Zwirgmaier



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