



7) If 0.51 moles of O_2 are used, how many moles of CH_4 are also used?

$$\frac{0.51 \text{ moles } \text{O}_2}{1} \times \frac{1 \text{ mole } \text{CH}_4}{2 \text{ mole } \text{O}_2}$$

Answer: 0.26 moles of CH_4

8) If 0.45 moles of H_2O are produced, how many moles of CO_2 are also produced?

$$\frac{0.45 \text{ moles } \text{H}_2\text{O}}{1} \times \frac{1 \text{ mole } \text{CO}_2}{2 \text{ mole } \text{H}_2\text{O}}$$

Answer: 0.23 moles of CO_2



9) Given 0.98 moles of N_2 , how many moles of NaN_3 are required?

$$\frac{0.98 \text{ moles } \text{N}_2}{1} \times \frac{2 \text{ moles } \text{NaN}_3}{3 \text{ mole } \text{N}_2}$$

Answer: 0.65 moles of NaN_3

10) Given 7.85 moles of NaN_3 , how many moles of Na are produced?

$$\frac{7.85 \text{ moles } \text{NaN}_3}{1} \times \frac{2 \text{ moles } \text{Na}}{2 \text{ mole } \text{NaN}_3}$$

Answer: 7.85 moles of Na



11) If 3.12 moles of Al_2S_3 are created, how many moles of S are also used?

$$\frac{3.12 \text{ moles } \text{Al}_2\text{S}_3}{1} \times \frac{3 \text{ moles } \text{S}}{1 \text{ mole } \text{Al}_2\text{S}_3}$$

Answer: 9.36 moles of S

12) If 0.32 moles of S are used, how many moles of Al are used?

$$\frac{0.32 \text{ moles } \text{S}}{1} \times \frac{2 \text{ moles } \text{Al}}{3 \text{ moles } \text{S}}$$

Answer: 0.21 moles of Al

Moles And Formulas Skill Practice 31

Ying-Ying Zheng



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