



# Hitachi Grinder Instructions

**Patrick Vollmar**



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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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web start with  $8.3 \text{ mole H}_2\text{S}$  and multiply it with  $\frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}}$  from the equation  $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$   $8.3 \text{ mole H}_2\text{S} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mole H}_2\text{O}} = 149.4 \text{ g H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.964 \text{ g/mol Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na  $2\text{C}_2\text{O}_4$  is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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ch<sub>3</sub>oh can be produced by the following reaction  
 $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$   
 calculate the theoretical yield of ch<sub>3</sub>oh if 68.5 g of co is reacted with 8.6 g of h<sub>2</sub>

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 $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$   
 which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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 $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$   
 a 80.0 grams of iodine v oxide i<sub>2</sub>o<sub>5</sub> reacts with 28.0 grams of carbon monoxide co determine the mass of iodine i<sub>2</sub> which could be produced  
 80 g i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub> xs 1 333.8 g i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub>o<sub>5</sub> 28 g co 1 mol co

**7 3 limiting reactant and percent yield problems** - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  
 text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3