

## **Molarity & Molality Notes and Practice**

Answer the questions below. **SHOW ALL WORK**, including units!! Watch your significant digits and **CIRCLE YOUR ANSWERS**.

### **Molarity**

Just a reminder, molarity is one of the many ways to measure concentration or the strength of a solution. When using molarity to measure concentration you must follow the formula below and then put a capital M at the end of your answer to let the world know you used the molarity formula.

$$M = \frac{\text{moles of solute}}{\text{Liters of solvent}}$$

1. Calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution.
2. What is the molarity of a solution containing 325 g of NaCl dissolved in 750. mL of solution?
3. 140 g of KCl is dissolved in 600. mL of water. What is the molarity?
4. 724.4 g of ammonium phosphate in 4500 mL of alcohol. What is the molarity of the solution?
5. You are making 2.2 L of 3.1 M silver nitrate solution. How many moles of solute are there?
6. How many grams of  $\text{MgCl}_2$  are needed to make 700 mL of a 1.4 M solution?
7. 93.2 g of copper (II) sulfate is mixed into 290. mL of water. What is the molarity?

### **Molality**

Molality is an additional way to measure the strength or concentration of a solution. It is abbreviated with a little m and is calculated only slightly differently than molarity. Here is the formula.

$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

# Molarity And Molality Practice Problems Worksheet

**Valery Ochkov**



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*General Chemistry Workbook* Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

**Student Study Guide Fundamentals of Chemistry**  
**by David E. Goldberg** Margaret G. Kimble,1994

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