

Concentration and Dilution Activity

You will be using a PHET simulation to explore the concept of molarity, and practice molarity and dilution calculations.

Purpose of this activity:

- Skills practiced: quantitative analysis of solutions
- Content reviewed: molarity and dilutions

Tasks:

- Complete the following activity.

Criteria for Success:

- Follow the directions below and complete the questions and calculations. Make sure to show your calculations and work for credit. Grade is based on completion, not accuracy.

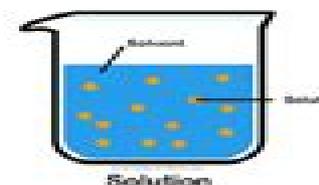
Simulation link:

[Concentration Simulation Link PHET Concentration](#)

Background information:

Much of chemistry takes place in **solution**. A solution has three parts:

- **Solute** – substance dissolved in another substance
- **Solvent** – the substance (usually liquid) in which the solute(s) is/are dissolved
- **Solute + Solvent = Solution**



An **aqueous** solution is a solution in which the solvent is water.

The most common unit of solution concentration is **molarity**. The definition of molarity is given below.

$$\text{Molarity (M)} = \frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$$

Our goal is to use a simulation to explore both **qualitative** and **quantitative** aspects of solutions. Follow the **instructions** below.

Open the simulation link. At the right of the screen, drag the Concentration sensor and place it in the initial solution. At the top right corner, change your solute to cobalt (II) nitrate. Shake the container 2 – 3 times to get a concentration reading of at least 0.300 mol/L.

First, let's examine some **qualitative properties** of solutions. We will examine how different actions affect the concentration of the solution.

1. Record your initial concentration.

It has a concentration of .300 mol/l , the color of the liquid is pink

2. Now we will add more solute. Shake some more cobalt (II) nitrate into the solution. Record your new concentration.

It has a concentration of 0.960 mol/ml, the liquid always gets a little darker.

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Jessica J Manson



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Holt Chemistry Ralph Thomas Myers,2004 **Chemistry** James E. Brady,John R. Holum,1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving *The Science Teacher's Toolbox* Tara C. Dale,Mandi S. White,2020-04-09 A winning educational formula of engaging lessons and powerful strategies for science teachers in numerous classroom settings The Teacher s Toolbox series is an innovative research based resource providing teachers with instructional strategies for students of all levels and abilities Each book in the collection focuses on a specific content area Clear concise guidance enables teachers to quickly integrate low prep high value lessons and strategies in their middle school and high school classrooms Every strategy follows a practical how to format established by the series editors The Science Teacher s Toolbox is a classroom tested resource offering hundreds of accessible student friendly lessons and strategies that can be implemented in a variety of educational settings Concise chapters fully explain the research basis necessary technology Next Generation Science Standards correlation and implementation of each lesson and strategy Favoring a hands on approach this book provides step by step instructions that help teachers to apply their new skills and knowledge in their classrooms immediately Lessons cover topics such as setting up labs conducting experiments using graphs analyzing data writing lab reports incorporating technology assessing student learning teaching all ability students and much more This book enables science teachers to Understand how each strategy works in the classroom and avoid common mistakes Promote culturally responsive classrooms Activate and enhance prior knowledge Bring fresh and engaging activities into the classroom and the science lab Written by respected authors and educators The Science Teacher s Toolbox Hundreds of Practical Ideas to Support Your Students is an invaluable aid for upper elementary middle school and high school science educators as well those in teacher education programs and staff development professionals **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Improving Student Comprehension of Stoichiometric Concepts** Connie Lynn Bannick Kemner,2007 **Scientific and Technical Aerospace Reports** ,1983 Chemical Principles Monty J. Montjar,1990 *Fertilizer Abstracts* ,1981 **Chemical Engineering Education** ,2002 **Physical Chemistry for the Life Sciences** Thomas Engel,Gary Drobny,Philip J. Reid,2008 KEY BENEFIT Physical Chemistry for the Life Sciences presents the core concepts of physical chemistry with mathematical rigor and conceptual clarity and develops the modern biological applications alongside the physical principles The traditional presentations of physical chemistry are augmented with material that makes these chemical ideas biologically relevant applying physical principles to the understanding of the complex problems of 21st century biology KEY TOPICS Physical Chemistry Biology MARKET For all readers interested in physical chemistry and biology **Science Citation Index** ,1992 Vols for 1964 have guides and journal lists **The Publishers' Trade List Annual** ,1985 *Kemia* ,1982

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