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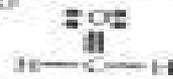
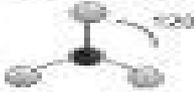
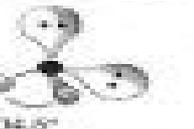
## Molecular Geometry

How can molecular shape be predicted using the VSEPR theory?

### Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The 3D shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

### Model 1 – Lewis Structures

	Lewis Structures		3-D Molecular Shape
1	$\text{H}_2\text{CO}$ 	$\text{H}_2\text{CO}$ 3 electron domains (3 bonding, 0 nonbonding)	 Trigonal Planar
2	$\text{BF}_3$ 	$\text{BF}_3$ 3 electron domains (3 bonding, 0 nonbonding)	 Linear
3	$\text{CH}_4$ 	$\text{CH}_4$ 4 electron domains (4 bonding, 0 nonbonding)	 Tetrahedral
4	$\text{NH}_3$ 	$\text{NH}_3$ 4 electron domains (3 bonding, 1 nonbonding)	 Pyramidal
5	$\text{H}_2\text{O}$ 	$\text{H}_2\text{O}$ 4 electron domains (2 bonding, 2 nonbonding)	 Bent
6	$\text{CO}_2$ 	$\text{CO}_2$ 2 electron domains (2 bonding, 0 nonbonding)	 Linear

Lone pair = ..

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