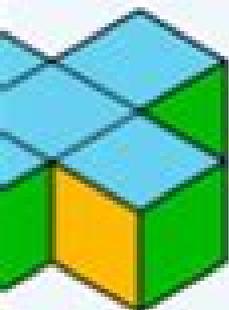


What factors affect the rate of reaction?

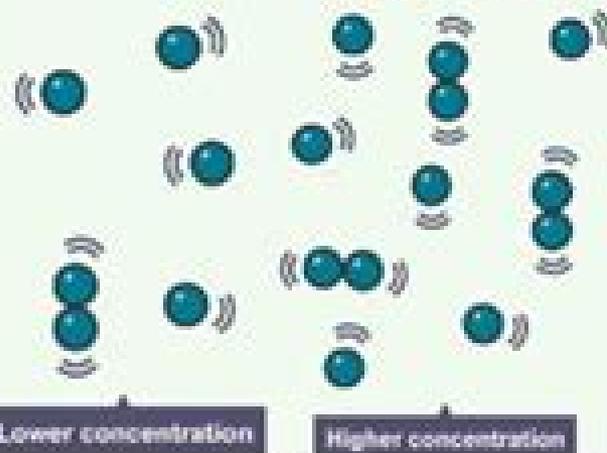
Surface Area



The **larger** the surface area of your reactants, the **faster** the rate of reaction

- More surface for a reaction so more reactant collisions.

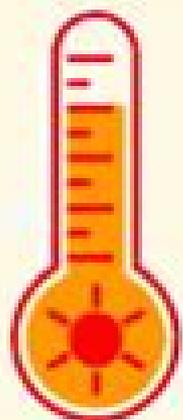
Concentration



The **higher** concentration of reactants, the **faster** the rate of reaction

- There are **more particles** in the same volume so there are **more collisions** between reactants.

Temperature



The **higher** the temperature of your reactants, the **faster** the rate of reaction

- Particles have **more energy** so move **faster**, increasing the rate of collisions.

Catalyst



A catalyst speeds up the reaction but is **NOT** used up

- Catalysts increase the number of successful collisions between reacting particles by **lowering activation energy**, thus reactions take place at **lower temperatures**.

Factors Affecting Reaction Rates Answers

Tobias Bleicker



Factors Affecting Reaction Rates Answers:

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