

F325 Acids, Bases and Buffers test ms

1. (a) completely dissociates/ionised (1) 2
 proton donor (1)
- (b) NO_3^- (1) 1
- (c) (i) $\text{pH} = -\log[\text{H}^+] / -\log(0.015)$ (1) = 1.82 / 1.8 (1) (Not 2) 2
 (ii) $[\text{H}^+] = 0.0075 \text{ mol dm}^{-3}$
 $\text{pH} = -\log(0.0075) = 2.12 / 2.1$ (1) 1

[6]

2. (a) (i) H^+ /proton donor (1) 1
 (ii) partially dissociates/ionises (1) 1
- (b) $\text{C}_6\text{H}_5\text{OH}(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{C}_6\text{H}_5\text{O}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
 acid 1 base 2 (1) base 1 acid 2 (1)
 1 mark for each acid-base pair
- (c) (i) $K_a = \frac{[\text{C}_6\text{H}_5\text{O}^-][\text{H}^+]}{[\text{C}_6\text{H}_5\text{OH}]}$ (1) 1

(ii) concentration = $38/94$ (1) = 0.40 mol dm^{-3} (1)
 (first mark for M_r of phenol – incorrect answer here will give ecf for remainder of question)

$$1.3 \times 10^{-10} = \frac{[\text{H}^+(\text{aq})]^2}{0.40} \quad (1)$$

('=' sign is acceptable)

$$[\text{H}^+] = \sqrt{(1.3 \times 10^{-10}) \times (0.40)} = 7.2 \times 10^{-6} \text{ mol dm}^{-3} \quad (1)$$

$$\text{pH} = -\log[\text{H}^+] = -\log 7.2 \times 10^{-6} = 5.14 \quad (1) \quad 5$$

3 marks: $[\text{H}^+]$ (1), pH expression (1), calc of pH from $[\text{H}^+]$ (1)

Common errors:

Without square root, answer = 10.28 (1)(1)(0)

Use of 38 as molar concentration does not score 1st 2 marks.

This gives an answer of 4.15 for 3 marks (1)(1)(1)

- (d) $[\text{H}^+(\text{aq})] = 1.99 \times 10^{-6} \text{ mol dm}^{-3}$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = K_a [\text{C}_6\text{H}_5\text{OH}(\text{aq})] / [\text{H}^+(\text{aq})]$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = 0.013 \text{ mol dm}^{-3}$ (1)

Calculation should use half the original concentration of phenol to find the concentration of sodium phenoxide in the buffer. This should then be doubled back up again.

Do not penalise an approach that uses the original concentration of phenol in the expression above.

[13]

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Vernon H. Naffier



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