

EQUILIBRIUM CONSTANT - WORKSHEET 2_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[NO_2]^2}{[N_2][O_2]}$$



$$K = \frac{[C_2H_6]}{[H_2]^3}$$



$$K = \frac{1}{[I_2(aq)]}$$

2. For the following reaction:



At equilibrium, $[Fe^{2+}] = 0.50 M$, $[Ag^+] = 1.0 M$, $[Fe^{3+}] = 1.50 M$

a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

b) If at equilibrium, $[Fe^{2+}] = 0.20 M$ and $[Ag^+] = 0.30 M$, what must be the concentration of Fe^{3+} ?

$$[Fe^{3+}] = 0.18 \text{ mol/L}$$



Calculate the equilibrium concentration of oxygen gas if the NO and NO_2 concentrations are equal at equilibrium.

$$[O_2] = 2.0 \times 10^{-6} \text{ mol/L}$$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of 1000°C , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-3}$$

Genetic Equilibrium Worksheet Answer Key

Julia Schneider



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Table of Contents Genetic Equilibrium Worksheet Answer Key

1. Understanding the eBook Genetic Equilibrium Worksheet Answer Key
 - The Rise of Digital Reading Genetic Equilibrium Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Genetic Equilibrium Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Genetic Equilibrium Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Genetic Equilibrium Worksheet Answer Key
 - Personalized Recommendations
 - Genetic Equilibrium Worksheet Answer Key User Reviews and Ratings
 - Genetic Equilibrium Worksheet Answer Key and Bestseller Lists
5. Accessing Genetic Equilibrium Worksheet Answer Key Free and Paid eBooks
 - Genetic Equilibrium Worksheet Answer Key Public Domain eBooks
 - Genetic Equilibrium Worksheet Answer Key eBook Subscription Services
 - Genetic Equilibrium Worksheet Answer Key Budget-Friendly Options

6. Navigating Genetic Equilibrium Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Genetic Equilibrium Worksheet Answer Key Compatibility with Devices
 - Genetic Equilibrium Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Genetic Equilibrium Worksheet Answer Key
 - Highlighting and Note-Taking Genetic Equilibrium Worksheet Answer Key
 - Interactive Elements Genetic Equilibrium Worksheet Answer Key
8. Staying Engaged with Genetic Equilibrium Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Genetic Equilibrium Worksheet Answer Key
9. Balancing eBooks and Physical Books Genetic Equilibrium Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Genetic Equilibrium Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Genetic Equilibrium Worksheet Answer Key
 - Setting Reading Goals Genetic Equilibrium Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Genetic Equilibrium Worksheet Answer Key
 - Fact-Checking eBook Content of Genetic Equilibrium Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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